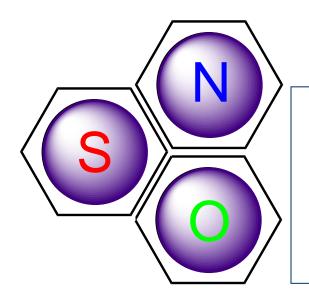


Heterocyclic Chemistry



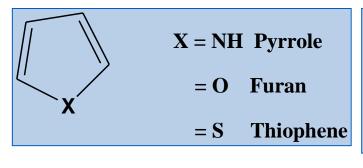
Five-membered Heterocycles

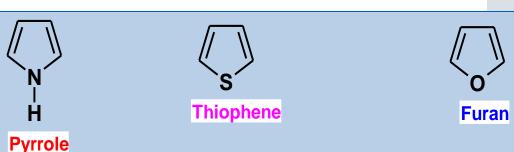
Pyrrole, Furan and Thiophene



A. Five-membered Rings with one Heteroatom

- **❖** The main reason for the study of pyrrole came from the work on the structure of haem; the blood respiratory pigment, and the chlorophyll; the green photosynthetic pigment of plants.
- **Thiophene and its derivatives occurs in <u>petroleum</u>**.
- **❖** Furan occurs widely in secondary plant metabolites, especially in terpenoids.
- **Unsubstituted pyrrole, furan, and thiophene are usually obtained from petroleum.**





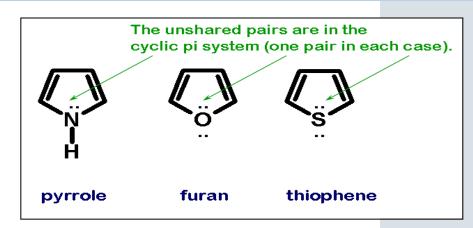


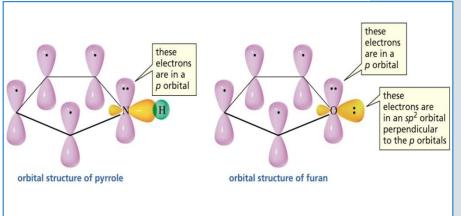
Structure and Aromaticity

- **Pyrrole furan and thiophene are aromatic because:**
- 1) they fulfill the criteria for aromaticity, the extent of delocalization of the nonbonding electron pair is decisive for the aromaticity, thus the grading of aromaticity is in the order of:

Furan < Pyrrole < Thiophene < Benzene this order is consistent with the order of electronegativity values for oxygen (3.44), nitrogen (3.04) and thiophene (2.56).

A. Five-membered Rings with one Heteroatom





The order of aromaticity

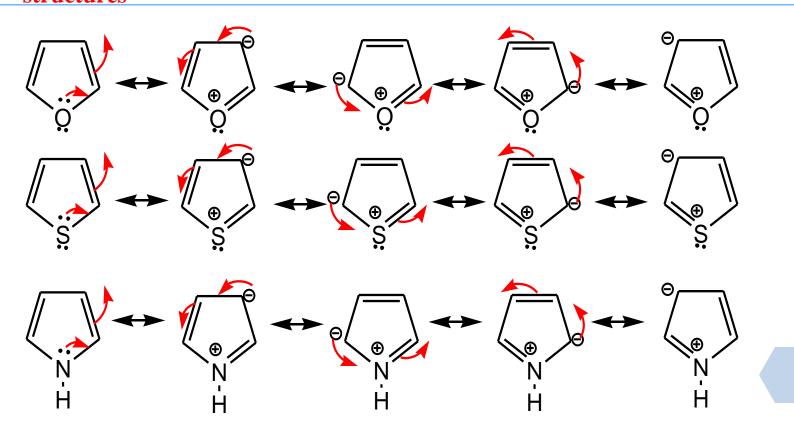
Benzene > Thiophene > Pyrrole > Furan



A. Five-membered Rings with one Heteroatom

Structure and Aromaticity

2) They tend to react by electrophilic substitution due appearance of –ve charge on carbon atoms due to delocalization as shown in the following resonance structures





A. Five-membered Rings with one Heteroatom

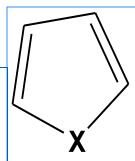
Structure and Aromaticity

3- Electrons not available for protonation—hence <u>not basic</u>

4- 6 π electrons over 5 ring atoms Electron rich... so more reactive than benzene towards electrophilic substitution.

The order of reactivity is:

Pyrrole > Furan > Thiophene > Benzene



X = NH Pyrrole

= O Furan

= S Thiophene





A. Five-membered Rings with one Heteroatom

Structure and Aromaticity

5- The pattern of reactivity with Electrophilic reagents.

Aromatic compounds By substitution addition followed by proton loss

[onium intermediate]

Order of reactivity: Pyrrole > Furan > Thiophene > Benzene

$$C_{2}\text{-attack} \qquad \begin{array}{c} & & & & \\ & &$$





Structure and Aromaticity

A. Five-membered Rings with one Heteroatom

NE

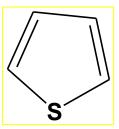
6-The order of aromaticity

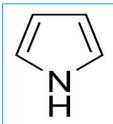
Benzene > Thiophene > Pyrrole > Furan

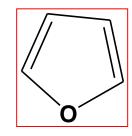
In case of Thiophene [S] donate & accept electrons..... so delocalization as complete as benzene

In case of Pyrrole [N] -Diene-like character CH₂=CH-CH=CH₂

In case of Furan [O] electronegativity more Diene-like character CH₂=CH-CH=CH₂











A. Five-membered Rings with one Heteroatom

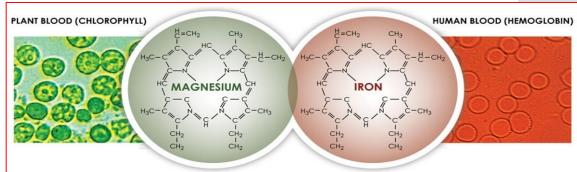
Sources & Synthesis

A) Sources

Pyrrole & Thiophene Coal Tar



Pyrrole ring
Porphyrin system....
Chlorophyll &
Hemoglobin



Furan Decarbonylation of Furfuraldehyde
Oat hulls, corn cobs or rice hulls



Oat hulls



corn cobs



rice hulls



B) Synthesis

A. Five-membered Rings with one Heteroatom

Thiophene

CH₃-CH₂-CH₂-CH₃ + 4 S
$$\xrightarrow{560}$$
 + 3 H₂S n- butane

Pyrrole

Furan

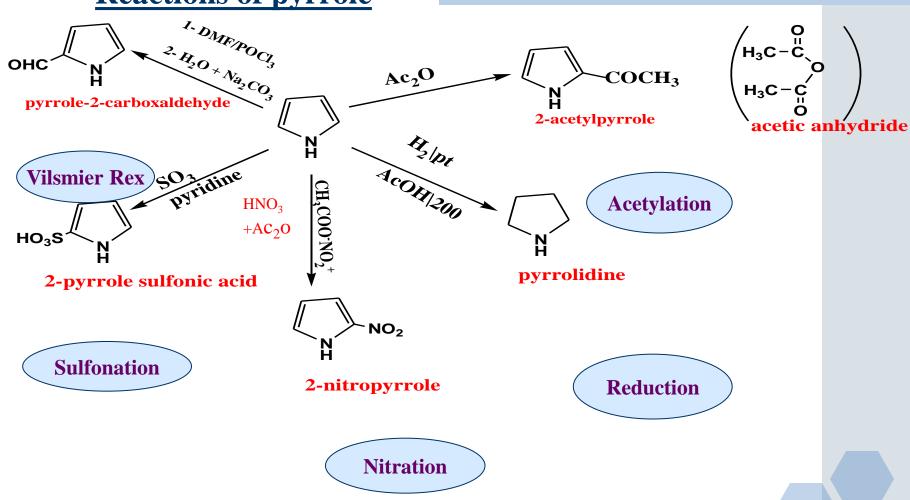
$$(C_5H_8O_4)_n \xrightarrow{H_2O,H^+} (CHOH)_3 \xrightarrow{-3H_2O} CHO$$

$$CH_2OH$$
Pentosan Pentose Furfural Furan



Reactions of pyrrole

A. Five-membered Rings with one Heteroatom





Reactions of Pyrrole

A. Five-membered Rings with one Heteroatom

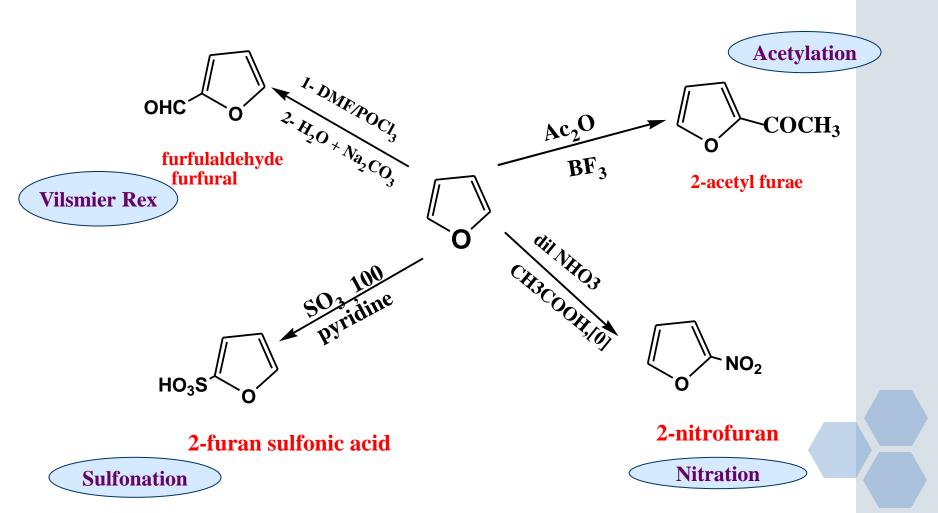
with carbene

Halogenation



Reactions of Furan

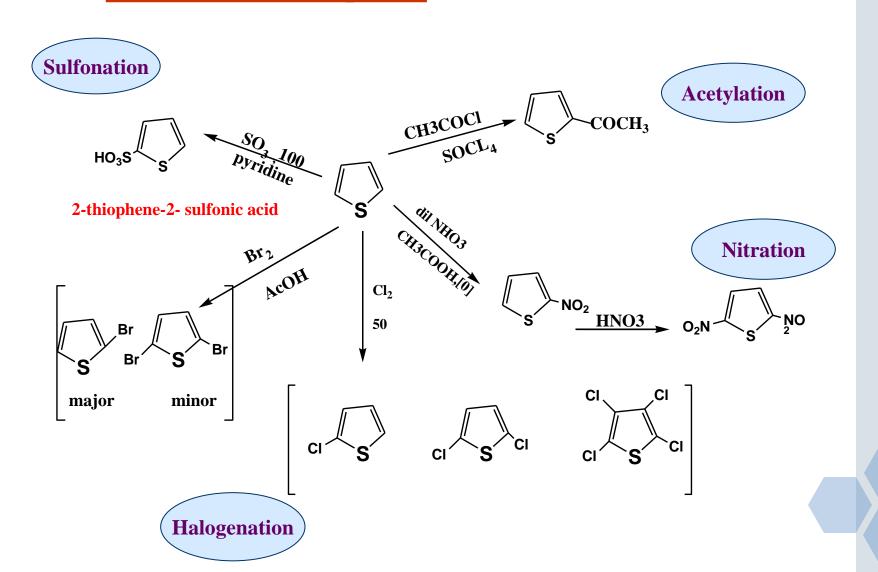
A. Five-membered Rings with one Heteroatom





Reactions of Thiophene

A. Five-membered Rings with one Heteroatom





A. Five-membered Rings with one Heteroatom

Five Membered Heteroaromatic Rings Containing 2X, at least one nitrogen



$$X = S$$
, o, or N



1,3-Azoles

Thiazole [1,3-thiazole]

Oxazole [1,3-oxazole]

Imidazole [1,3-diazole]

1,2-Azoles

Isothiazole [1,2-thiazole]

Isoxazole [1,2-oxazole]

Pyrazole [1,2-diazole]



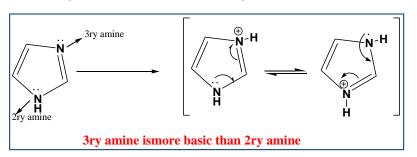
A. Five-membered Rings with one Heteroatom

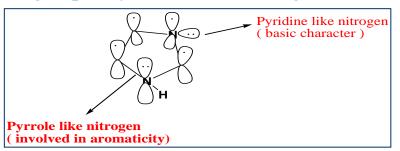
Five Membered Heteroaromatic Rings Containing 2X, at least one nitrogen

Aromaticity & Basicity



Strong base, due to the greater electron releasing capacity of the two nitrogen's

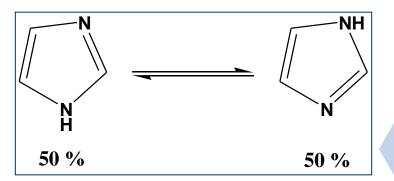




Importance of the ring

Building blocks as Histidine and Histamine

It exist in two tautomeric forms

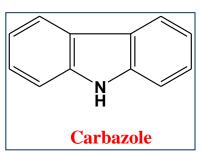




A. Five-membered Rings with one Heteroatom

Fused Five Membered Heteroaromatic Rings With one X





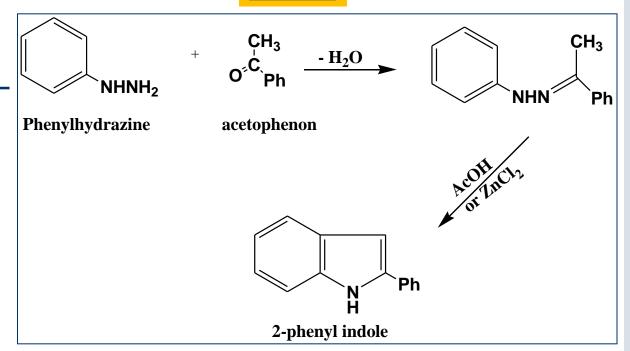


Fused Five Membered Heteroaromatic Rings With one X

A. Five-membered Rings with one Heteroatom

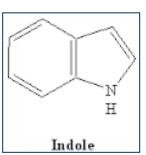
Indole

Fischer Synthesis



Basicity of Indole

Not basic because lone pair is delocalized and contributed to the Aromatic system





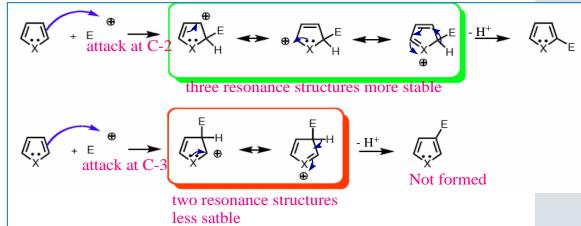


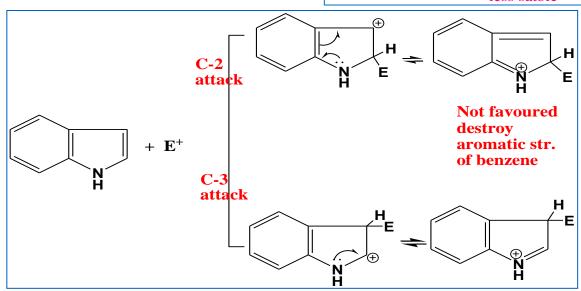
Fused Five Membered Heteroaromatic Rings With one X

A. Five-membered Rings with one Heteroatom

Indole

<u>It undrego Electrophilic</u> <u>Subistitution at Position 3</u>







A. Five-membered Rings with one Heteroatom



Indole-3-carboxylaldehyde

quinoline







A. Five-membered Rings with one Heteroatom

Fused Five Membered Hetero aromatic Rings With two X

Benzimidazole ...

In Vit B12: Commercially as an parasiticide





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