



Al-Mustaqbal University College

Department of Medical Physics

First Class

Organic Chemistry

Lec 5 Alkynes

MSc .Doaa Nassr

MSc .Issa Farahan

Alkynes

- Members of the alkyne group have a triple bond between two carbon atoms.
- Two hydrogen atoms have been removed from each of two adjacent carbon atoms, thereby allowing the two adjacent carbon atoms to form a triple bond
- General formula is C_nH_{2n-2}

- Begins with ethyne (acetylene)



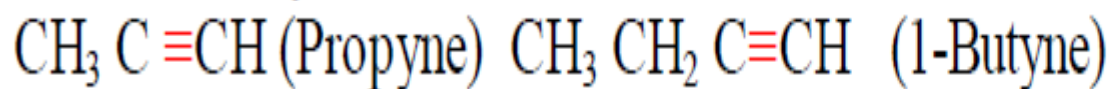
Ethyne (acetylene)

- C_2H_2

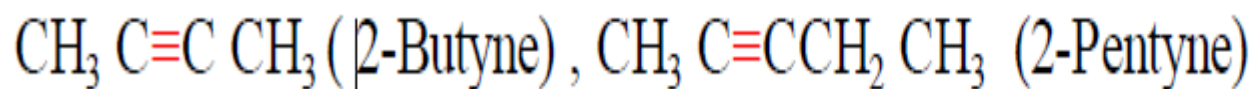
Classification of alkyne :-

Alkyne are further classified as terminal or non-terminal alkynes according as the triple bond is present at the carbon chain or within the carbon chain.

Terminal alkynes



Non-Terminal alkynes



IUPAC Rules for Alkyne Nomenclature:-

The IUPAC rules for naming alkynes are 1)-the same as those for alkenes except that the ending is -yne .

2. The **yne** suffix (ending) indicates an alkyne or cycloalkyne.

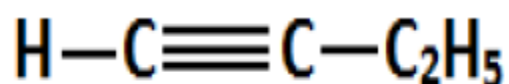
3. The longest chain chosen for the root name must include both carbon atoms of the triple bond.

4. The root chain must be numbered from the end nearest a triple bond carbon atom. If the triple bond is in the center of the chain, the nearest substituent rule is used to determine the end where numbering starts.

5. The smaller of the two numbers designating the carbon atoms of the triple bond is used as the triple bond locator.

5. If several multiple bonds are present, each must be assigned a locator number.

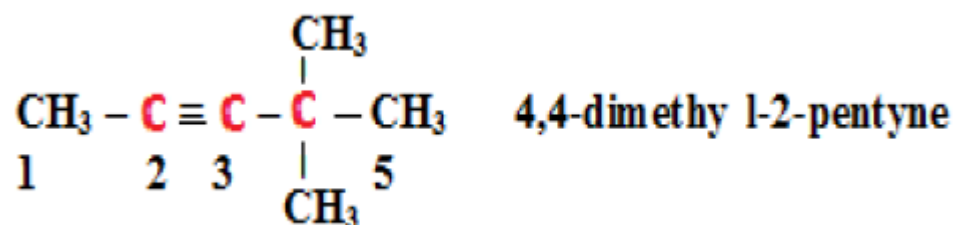
The following examples illustrate the rules:



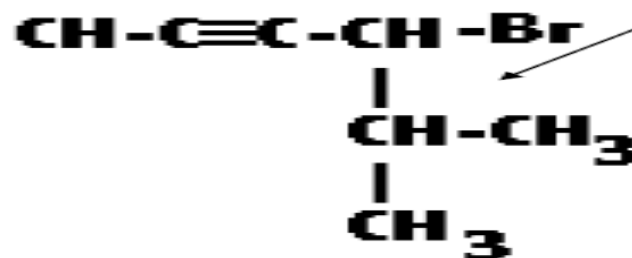
1-butyne



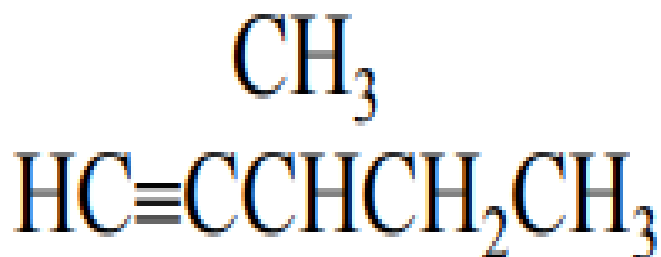
2-butyne



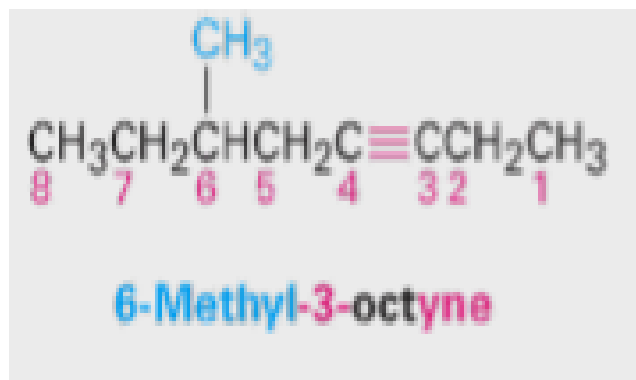
$\text{CH}_3\text{C}\equiv\text{CH}$ methylacetylene (propyne)



4bromo-5-methyl-2-hexyne



3-methyl-1-pentyne

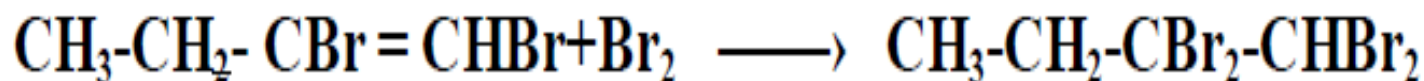
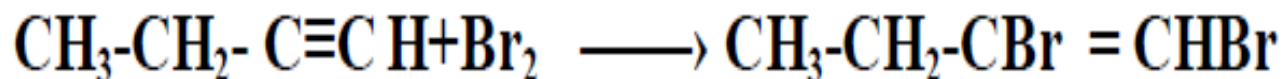


6-Methyl-3-octyne

Reaction :-

Most of the reactions of alkynes are similar to those of alkenes. The same reagents that add to carbon-carbon double bond also add to carbon-carbon triple bond . But it is possible to add two molecules of reagent to each alkyne.

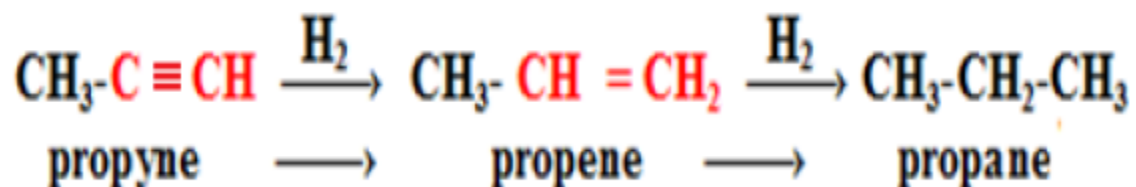
1- Addition of Halogens (Halogenation) :-



1-butyne \longrightarrow 1,2-dibromo-1-butyne \longrightarrow 1,1,2,2-tetrabromobutane

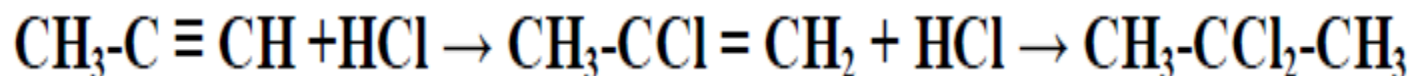
2-Addition of Dihydrogen (Hydrogenation)

The addition of H₂ to alkyne is obtained by adding hydrogen gas to alkyne with the use of metal (Ni, Pd, ...) catalysis to give alkene in the first step and an alkane in the second step



3- Addition of Halogen halide (Hydrohalogenation) :-

The addition of HX is obtained according to Markovnikov's Rule (the acid hydrogen (H) gets attached to the carbon with more hydrogen substituents, and the halide (X) group gets attached to the carbon with more alkyl substituents) (Markovnikov's Rule)



4- Addition of Water(Hydration) :-

One difference between the acid catalyzed hydration of alkenes and that of alkynes. Alkenes form alcohol. Alkynes form compounds containing C=O bond.

