

Al-Mustaqbal University College

Practical organic chemistry

The second stage

IDENTIFICATION OF ALCOHOLS

Submitted by

Assistant Lect: Zahraa falah

Alcohols

Alcohols are organic compounds that may be considered as derivatives of water in which one of the hydrogen atoms of a water molecule (H-O-H) has been replaced by an alkyl or substituted alkyl group. Therefore, properties of alcohols may be related to the properties of both water and hydrocarbons. The alkyl group could be primary, secondary, or tertiary, and maybe an open-chain or cyclic. Accordingly, alcohols may be defined as organic compounds that contain hydroxyl groups attached to an alkyl, substituted alkyl, or a cyclic alkyl group

Nomenclature of Alcohols

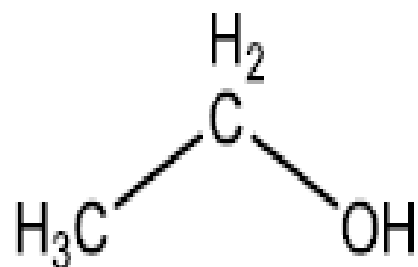
1- Name the longest carbon chain that contains the carbon atom bearing the —OH group. Drop the final -e from the alkane name, and add the suffix -ol.

2- Number the longest carbon chain starting at the end nearest the —OH group, and use the appropriate number, if necessary, to indicate the position of the —OH group.

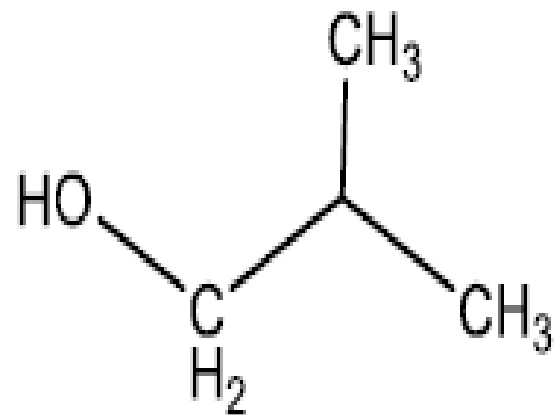
3- Name the substituents, and give their numbers as for an alkane or alkene.



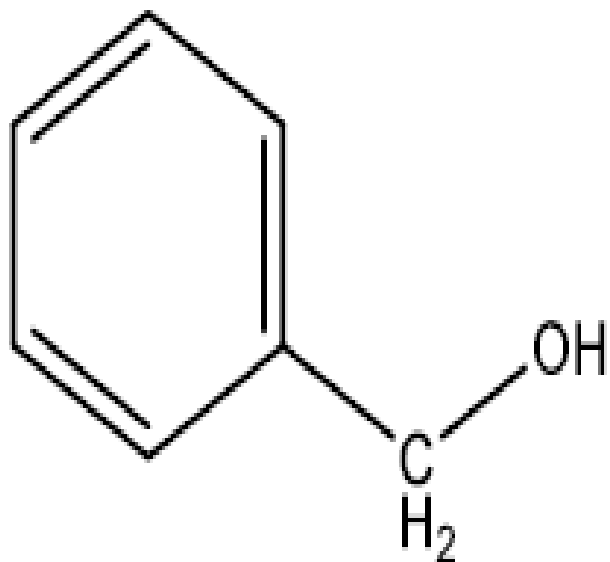
Methanol



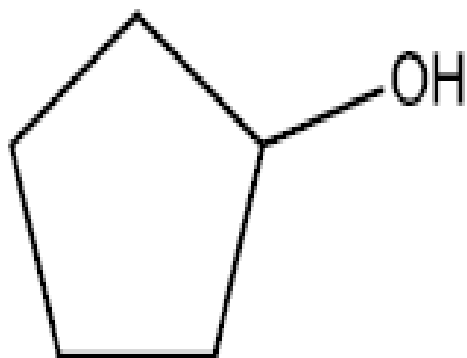
Ethanol



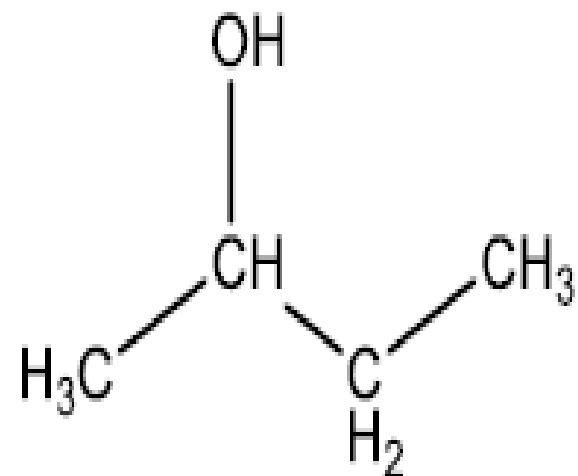
Isobutanol



benzyl alcohol



cyclopentanol



sec-butanol

Physical Properties:

- 1- Alcohols are colorless liquids with a special faint odor. Benzyl alcohol and cyclohexanol have characteristic odors.*
- 2- Aliphatic alcohols burn with a blue flame (without smoke) while aromatic alcohols burn with a yellow smoky flame.*
- 3- Boiling points of alcohols are considerably high (being associated liquids); they increase as the molecular weight (number of carbons) increases.*
- 4- Alcohols are miscible with water except for benzyl alcohol, cyclohexanol, and sec-butanol (which is very slightly soluble in water).*

Solubility Classification:

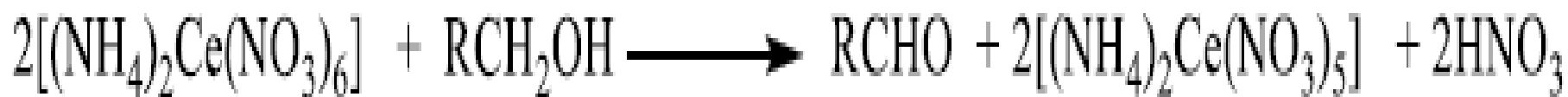
Alcohols are polar compounds because of the presence of the hydroxyl group which is also responsible for their ability to form hydrogen bonding. The degree of the polarity depends on the size of the alkyl side chain; the polarity decreases as the size of the alkyl side chain increases, or in another word, as the hydroxyl group hydrocarbon ratio of alcohols increases, their water solubility increases, and vice versa. Besides, low molecular weight alcohols are soluble in water due to hydrogen bonding ability with water molecules. Therefore, alcohols that are soluble in water and ether are classified under class S1 such as ethanol and methanol. Alcohols that are insoluble in water are related to class N such as benzyl alcohol, sec-butanol, and cyclohexanol.

Chemical Properties:

- 1- Alcohols are neutral compounds that don't change the color of*
- 2- litmus paper.*
- 3- All reactions of alcohols are related to its active hydroxyl group and are of two types:*
 - a) removal of the hydroxyl itself as in the reaction with hydrogen halides to form alkyl halides or in the dehydration reaction to form a double bond.*
 - b) removal of the proton only from the active hydroxyl as in the formation of esters or in the reaction with active metals such as sodium.*

1. General test (Ceric ammonium nitrate reagent)

Ceric ammonium nitrate (yellow solution) is an oxidizing agent that reacts with alcohols to give a red complex and with phenols to give a brown to greenish brown precipitate. Each mole of the alcohol requires two moles of the reagent. The red colored complex is an intermediate for the oxidation of alcohols by the Ce (IV) solution. This red color disappears after a reasonable time due to completing the oxidation of this intermediate and the reduction to the colorless Ce (III) solution producing the corresponding aldehyde or ketone.



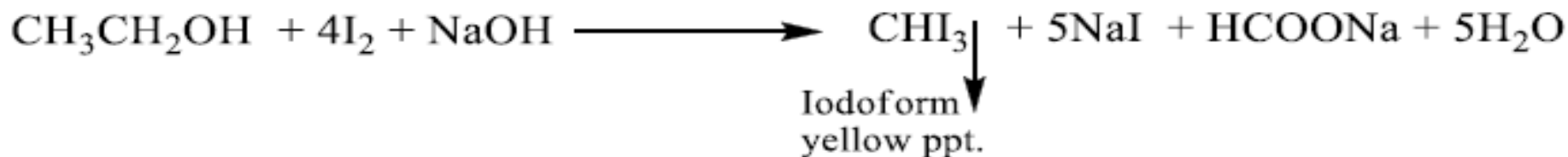
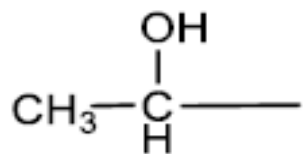
Procedure:

- 1- Water soluble (miscible) alcohols; mix two drops of the alcohol with one drop of ceric ammonium nitrate solution. A red complex indicates a positive test.*
- 2- Water-insoluble (immiscible) alcohols; mix two drops of the alcohol with 0.5 ml dioxane, shake well, and add one drop of the reagent to get a positive red complex.*
- 3- This test gives positive results with primary, secondary, and tertiary alcohols (up to 10 carbons), polyhydroxylated compounds such as carbohydrates, and hydroxylated carboxylic acids, aldehydes and ketones.*

Specific tests

a) Iodoform (Haloform) test

This test is specific for alcohols which have a free methyl group and a hydrogen attached to the carbon bearing the hydroxyl group such as ethanol and sec-butanol.



The overall reaction is:

The alcohol is oxidized to the corresponding aldehyde or ketone by the action of the produced oxidizing agent 'sodium hypoiodite', which also causes the aldehyde or ketone to be tri-iodinated on the terminal methyl group producing iodoform as a yellow precipitate.

Procedure:

Dissolve about 3 drops of the alcohol in about 2 mL of distilled water (or 2 mL of dioxane for water insoluble compounds), add about 1 ml of 10% sodium hydroxide solution, then add iodine solution dropwise with shaking until either a yellow iodoform precipitate is formed, in which case the test is positive and is completed, or the dark color of the iodine solution is present. In the latter case allow the solution to stand for 3 minutes during which period check for the appearance of the yellow precipitate at the bottom of the test tube. If there is no precipitate, warm the solution in a water bath (60 °C) for about 3 minutes with shaking from time to time and check for

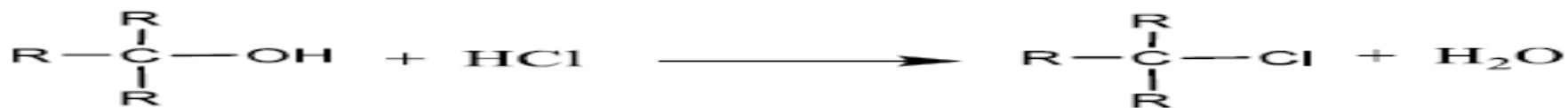
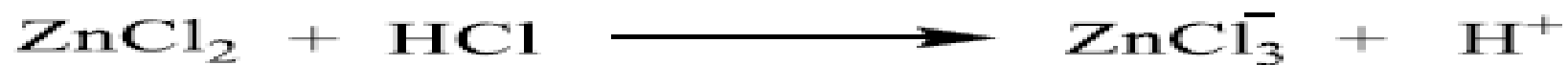
the yellow precipitate. During warming, if the color of iodine disappears, add few additional drops of iodine solution with shaking until either the yellow precipitate is formed or the dark iodine color persists, and then complete warming. Then get rid of the excess iodine by the addition of 10% sodium hydroxide solution dropwise with shaking to obtain the yellow precipitate. If the precipitate is not formed, allow the solution to stand for 10 minutes to get the positive result. Finally, if no precipitate is formed after the 10 minutes-standing period, dilute the solution with an equal volume of

distilled water to obtain the iodoform precipitate. It is important to proceed through all these steps so that only at the final step you can say that the test is negative. Both ethanol and sec-butanol give positive iodoform test and they can be differentiated only by testing their solubility in water; sec-butanol is less soluble in water than ethanol.

b) Lucas test

This test often provides classification information on alcohols and is used to distinguish between the different types of alcohols (primary, secondary, or tertiary). It depends on the formation of alkyl chloride as a second liquid phase. Lucas reagent is prepared from anhydrous zinc chloride and concentrated hydrochloric acid. Zinc chloride is added to increase the ionization of hydrochloric acid. Benzyl alcohol shows the fastest positive result. Tertiary alcohols are faster in the formation of conjugated halides than secondary

alcohols. Primary alcohols and methanol don't react and don't form two layers.



Procedure

Mix 2-4 drops of the alcohol with few drops of Lucas reagent and observe the results:

- 1. benzyl alcohol gives the immediate result as shown by the appearance of two phases.*
- 2. tertiary alcohols give two phases that separate within 2-3 minutes.*
- 3. secondary alcohols give two phases that separate after 15-20 minutes (giving a cloudy solution).*
- 4. In primary alcohols, one layer appears*

d) Oxidation reaction (with potassium permanganate)

Primary alcohols oxidase to cross ponding carboxylic acid, Secondary alcohols oxidase to a cross ponding ketones, and tertiary alcohols don't oxidase the result from oxidation of primary alcohols rapidly oxidase to form carboxylic acids with manganese dioxide(brown ppt.) ,Secondary alcohols oxidase after 3-5 min., and tertiary alcohols don't oxidase (the color of solution is purple).

d) Oxidation reaction (with potassium permanganate)

Procedure

To 0.5ml of alcohol add 4drops of acetic acid, and 5-8 drops from potassium permanganate and then observed the result.